



Uranium Removal from Aqueous Solution Using Ion-exchange Resin DOWEX[®] 2x8 in the Presence of Sulfate Anions

M. R. Kosari^a, H. Sepehrian^{*b}, R. Salamat Ahangari^c

^a Department of Energy Engineering, Sharif University of Technology, Azadi Avenue, Tehran, Iran.

^b Nuclear Science and Technology Research Institute, End of North Karegar Avenue, Tehran, Iran

^c Department of Chemistry, Faculty of Basic Science, Azarbaijan Shahid Madani University, Tabriz, Iran

PAPER INFO

Paper history:

Received 01 June 2016

Received in revised form 04 October 2016

Accepted 11 November 2016

Keywords:

Dowex 2x8

Anion Exchanger Resin

Anion Effect

Adsorption

Uranium

Heavy Metals

ABSTRACT

The current study was aimed at uranium (VI) ions adsorption using a anion exchanger resin, Dowex[®]2x8, in the presence of sulfate anions. Evaluation of operational parameters including contact time, pH, initial concentration of uranium ions, and presence of various anions (including phosphate, sulfate, chloride, fluoride, and nitrate) in the solution was assessed on sorption performance of Dowex 2x8. Employing the Langmuir and Freundlich binding models, isotherm analysis was performed to clarify the relation between uranium ions adsorbed on the resin and the concentration of remaining uranium ions in the aqueous phase. To reveal the rate-limiting step of the sorption process, kinetic modeling using pseudo first-order and pseudo second-order equations was also carried out. Maximum sorption capacity for Dowex 2x8 anionic exchanger resin on uranium removal process was found to be 303.03 mmol g⁻¹. It was also disclosed that existence of diverse anion contents in the solution could affect the uranium sorption using Dowex 2x8 anion exchanger resin.

doi: 10.5829/idosi.ije.2016.29.12c.06

NOMENCLATURE

q	Sorption capacity of the adsorbent (mmol g ⁻¹)	C _e	Equilibrium concentration of U(VI) ions in solution (mmol l ⁻¹)
C _i	Initial concentration of uranium ions (mmol l ⁻¹)	k ₁	Langmuir isotherm constant (l mmol ⁻¹)
C _f	Final concentration of uranium ions (mmol l ⁻¹)	n	Freundlich constants corresponded to the sorption intensity
V	Volume of uranium solution (ml)	k _f	Freundlich constants corresponded to the sorption capacity
m	Mass of adsorbent (g)	q _t	Amount of the adsorbed uranium (mmol g ⁻¹) on adsorbent at time t
q _e	Amount of uranium ions adsorbed per unit weight of the sorbent (mmol g ⁻¹)	K ₁	Pseudo first-order rate constant of adsorption (l min ⁻¹)
q _{max}	Maximum sorption capacity (mmol g ⁻¹)	K ₂	Pseudo second-order rate constant of adsorption (g mmol ⁻¹ min ⁻¹)
		q _m	Monolayer sorption capacity

1. INTRODUCTION

Recently, uranium, the heaviest and radioactive element among rare-abundant elements in the earth's surface, has been found to have essential role for supplying the required energy of human life. Two important steps for

extraction of uranium from its minerals including mining and milling processes have propagated great deals of processed and unprocessed uranium into the environment with both natural and anthropological events.

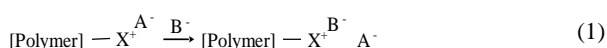
Having two kinds of toxicity including chemical and radioactive, uranium can severely affect human body. After uranium entrance to the body, it will mix with blood and absorb by different tissues like kidneys,

*Corresponding Author's Email: hsepehrian@aeoi.org.ir (H. Sepehrian)

bones and skin, in which it can cause different cancers due to its radioactivity. Inhalation of insoluble uranium compounds is more dangerous because of its accumulation in respiratory tract and having no tendency to leave [1].

As for removal of toxic metals from contaminated media, so far various methods and materials have been investigated [2-7]. However, most of those methods for complete removal of heavy metals are inefficient due to their high costs, technical problems, and production of toxic chemical sludges. Among separation techniques, adsorption due to its merits such as low cost, easy operation and high efficiency has deserved special attention toward the removal of heavy metals [8]. On the other hand, grate deal of attempts was made by researchers for use of new adsorbents. Diverse group of synthetic and natural adsorbents were developed and used to reach the highest efficacy in adsorption of heavy metal ions, but still, most of researchers are seeking to find new adsorbents having considerable advantages such as maximum adsorption capacity, low cost and regeneration ability.

Metal ions adsorption mechanism on the surface or on the active sites of the adsorbents can be related to the nature of sorbents. Recently, many reports have been focused on the use of ion exchange resins including strong and weak acidic/basic anion or cation exchanger resins as sorbents [9-11]. The interchange of ions between two phases happens in an ion exchange process. When a resin- a cross-linked polymer network-consisting of ion to which is electrostatically bound, encounters with a solution including ions of the same charge, an ion exchange can occur that the extent of which depends on the concentration of the ion(s) in the solution [12]. The mechanism for an anions exchange between an insoluble phase and a solution phase can be depicted as below [12] (Equation (1)):



For the removal of macromolecules such as uranium, so far many techniques have been proposed and uranium removal in different conditions was investigated. There are various specific and general sorbents, which can adsorb uranium singly or in a matrix. Iminodiacetic acid, namely IDA, is a specific sorbent for uranium removal that its uranium adsorption just effectively occur around pH=2.5 [13]. In addition, graphene oxide (GO) is a general sorbent with the ability of adsorption of uranium along with other ions through the aqueous solutions [14]. However, there are operational problems through industrial application such as interference of uranium with other anions for sorbents surface or active sites of the sorbent. As a matter of fact, it was confirmed that existence of different anion contents whether single or binary could adversely influence on uranium removal using various sorbents or ion exchanger resins.

Currently, uranium (VI) ions sorption using an anion exchanger resin, namely Dowex 2x8, was investigated as a function of contact time, initial concentration, pH and diverse anions in the solution. For better understanding of the sorption process, a mathematical description by applying isotherm and kinetic models on experimental data was eventually performed.

2. EXPERIMENTAL

2. 1. Materials

Uranium nitrate ($\text{UO}_2(\text{NO}_3)_2 \cdot 6\text{H}_2\text{O}$) used for the sorption experiments was of analytical grade with 99.99% assay. Sodium sulfate ($\text{Na}_2(\text{SO}_4)$) was also used in the current study as phosphate anion source. Dowex 2x8, a strongly basic anionic exchanger resin, 200-400 mesh was obtained from Sigma-Aldrich (Milwaukee, WI, USA). Other chemicals and reagents used were of analytical grade from Merck (Darmstadt, Germany).

2. 2. Sorption Experiments

Batch sorption experiments were implemented to evaluate sorption properties of Dowex 2x8 anion exchanger resin for uptake of uranium ions from aqueous solution. In a typical test, 20 mg of adsorbent were added into 10 ml of U(VI) solution whose initial concentration was situated on a certain value. Additionally, sulfate anions with specific concentration (2655 mg L^{-1}) existed in the solution. After a desired time of stirring and attaining equilibrium, adsorbent was segregated from solution by filtration. Consequently, the remained solution was analyzed by an inductively coupled plasma-atomic emission spectrometry (ICP-AES), for its uranium content. The sorption capacity (q , mg g^{-1}) and sorption percentage (uptake, %) were calculated using following equations (Equations (2) and (3)):

$$q(\text{mg g}^{-1}) = (C_i - C_f) * \frac{V}{m} \quad (2)$$

$$\% \text{uptake} = \left(\frac{C_i - C_f}{C_i} \right) * 100 \quad (3)$$

Working solution pH effect on the removal behavior of Dowex 2x8 was determined at a constant temperature of 25°C for 2 h. In batch tests, 20 mg of the adsorbent was equilibrated with 10 ml of the solution containing 200 mg L^{-1} of U(VI) ions at various pH quantities. The adsorption of uranium was investigated in the pH range of 2-9.

Isotherm analysis was performed for assessment of the relationship between the amount of uranium ions adsorbed on Dowex 2x8 surface and concentration of remaining metal ions in solution. Out of several binding models, the Freundlich and Langmuir isotherms have been extendedly used to describe the experimental results of sorption isotherm. The aforementioned isotherm models are mathematically represented in linear forms as (Equations (4)-(6)):

Langmuir isotherm [15]:

$$\frac{C_e}{q_e} = \frac{1}{q_{\max}k_1} + \frac{C_e}{q_{\max}} \quad \text{Linear form} \quad (4)$$

Freundlich isotherm [16]:

$$q_e = k_f C_e^{1/n} \quad (5)$$

$$\log q_e = \log k_f + \frac{1}{n} \log C_e \quad \text{Linear form} \quad (6)$$

Kinetics analysis was carried out to reveal the rate-limiting step of sorption process such as chemical reaction or mass transfer. On that account, two kinetics models, namely pseudo first-order and second-order, are applied on experimental data of uranium sorption onto Dowex 2x8. The pseudo first-order and second-order formulas are expressed by (Equations (7) and (8)) [17]:

Pseudo first-order model:

$$\log \left(\frac{q_e}{q_e - q_t} \right) = \left(\frac{K_1}{2.303} \right) t \quad (7)$$

Pseudo second-order model:

$$\frac{t}{q_t} = \frac{1}{K_2 q_e^2} + \frac{t}{q_e} \quad (8)$$

The solution resemblance to real effluent was employed to assess the anion effect on sorption performance of Dowex 2x8 anion exchanger resin. For this purpose, adsorption properties of the resin were tested during the presence of various anion contents in the solution with a concentration range of those anions.

3. RESULTS AND DISCUSSION

3. 1. Effect of pH Performing pH tests on uranium removal, sorption efficiency of uranium ions versus pH values was sketched in Figure 1. The result shows that Dowex 2x8 anion exchanger resin is capable of uptaking U(VI) ions in pH ranges 2 to 9 effectively. This behavior can be ascribed to capacity of active sites existing on resin's surface for holding U(VI) ions and their affinity towards various uranium anionic species.

As can be seen, as pH quantity increases, the amount of H^+ ions decreases and uranium ions uptake is gently reached to maximum sorption efficiency and is remained at a plateau after pH 3. Achieving the high efficiency may be by virtue of the formation of different complexes of uranium in the presence of hydroxyl ions and sulfate anions in the medium. Thus, sorption of U(VI) species at $pH > 2$ than $pH = 2$ enhances because all of uranium dioxide cations in the form of UO_2^{2+} at $pH < 2$ change to anionic species such as $UO_2SO_4^{2-}$ and $(UO_2)OH_4^{2-}$ at higher pHs whose affinity towards Dowex 2x8 anion exchanger resin are huge.

3. 2. Isotherm Study Adsorption rate of uranium ions onto the resin showed a striking variation with the initial concentration of uranium ions (Figure 2a). It is

obvious that adsorption efficiency increased with increasing the initial concentration of uranium (VI) ions. The Langmuir and Freundlich isotherms were applied to experimental data. The Langmuir sorption isotherm model is based on the assumption that all the sorption sites are energetically identical and sorption occurs on a structurally homogeneous sorbent [18]. While, the Freundlich isotherm model assumes that enthalpy of adsorption is independent of the amount adsorbed and heterogeneous surface is involved in the sorption process [19].

Concerning with Figure 2b, parameters of Langmuir model can be calculated from the trend line equation. Based on Freundlich isotherm data (see Figure 2c), by plotting $\log q_e$ vs. $\log C_e$, amounts of K_f and n parameters can be evaluated from the intercept and slope, respectively. An analogous between Langmuir and Freundlich isotherms indicates that the former model can satisfactorily fit experimental data. It can be concluded that the sorption process is affected by the formation of monolayer coverage of the sorbate (uranium) on the homogenous sorbent surface. The isotherm constants given in Table 1 show good agreement between experimental data and the Freundlich isotherm result based on regression value. The maximum adsorption capacity (q_{\max}) of the Dowex 2x8 anion exchanger resin was found to be $303.03 \text{ mmol g}^{-1}$.

3. 3. Kinetics Study

To perceive the relation between the contact time and sorption performance of Dowex 2x8 anion exchanger resin, equilibrium experimental data of uranium (VI) ions were measured at different time intervals up to 24 h.

Regarding the profile illustrated in Figure 3a, the amount of uranium (VI) ions sorption efficiency of the adsorbent increased with time and equilibrium condition was achieved after 2 h.

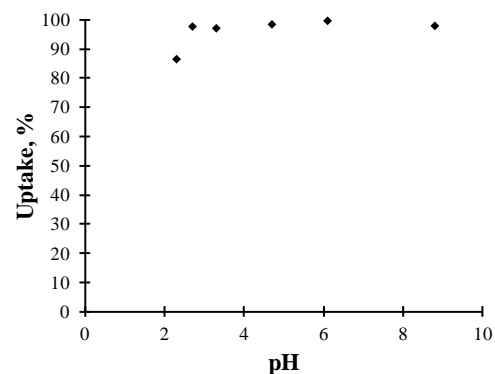


Figure 1. Effect of working solution pH on uranium (VI) adsorption using Dowex 2x8 anion exchanger resin
Conditions: $C_0(U)$: 5 mg.l^{-1} , sulfate concentration: 2655 mg.l^{-1} , agitation time: 2 h, and temperature: 25°C

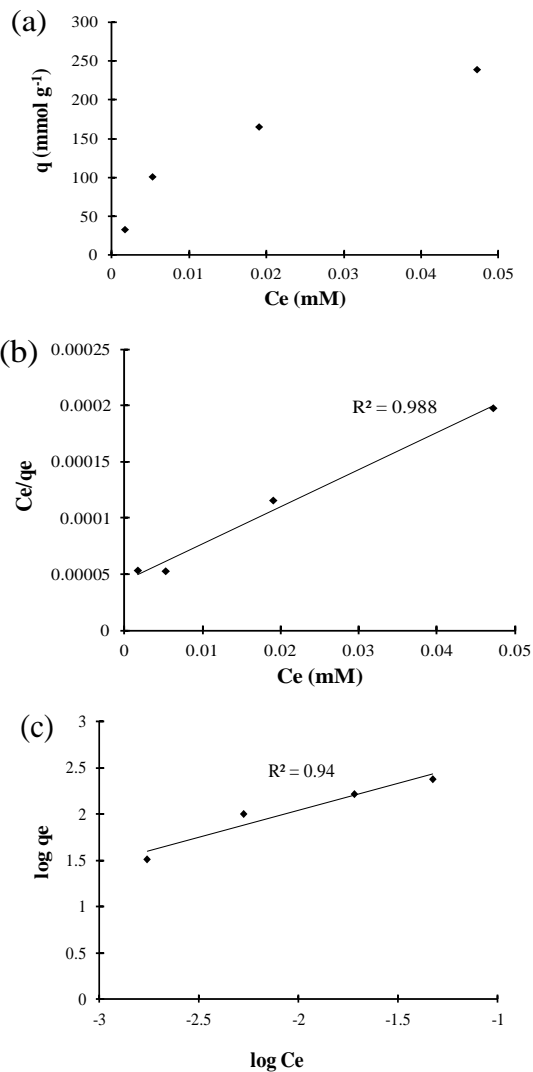


Figure 2. (a) Effect of initial concentration of uranium ions on Dowex 2x8 sorption behavior, (b) Langmuir adsorption isotherm curve of U(VI) ions uptake and (c) Freundlich adsorption isotherm curve of U(VI) ions uptake. Conditions: initial pH value: 3 sulfate concentration= 2655 mg.l⁻¹, agitation time: 2 h, and temperature= 25°C.

TABLE 1. Isotherm constants for uranium (VI) adsorption onto Dowex 2x8 anion exchanger resin

Isotherm		Parameters	
Langmuir equation	q_{max} (mmol g ⁻¹)	K_L (L mM ⁻¹)	R^2
	303.03	82.5	0.988
Freundlich equation	n	K_f (mmol g ⁻¹ mM ⁻¹)	R^2
	1.72	1587.44	0.94

On account of investigating the governing mechanism on U(VI) ions sorption, the rate of uranium ion sorption, one of the important traits that portrays

sorption efficacy, was determined by applying the pseudo first-order and pseudo second-order kinetic models (Table 2).

The pseudo first-order kinetic model has been used for the description of reversible equilibrium between liquid and solid phase, whereas, the second-order kinetic model assumes that the rate limiting step may be chemical adsorption [20]. An examination of kinetics outcome shows that pseudo second-order model fitted the kinetic data of the Dowex 2x8 anion exchanger resin better than that of pseudo first-order. The affirmation of pseudo second-order model indicates that the concentration of both sorbate (uranium) and sorbent (Dowex 2x8) are involved in the rate-limiting step of the sorption process [21].

3. 4. Effect of Various Anions on Uranium Removal

One factor that leads to debilitation of most sorbent's effectiveness in uranium sorption process is the presence of interfering ions.

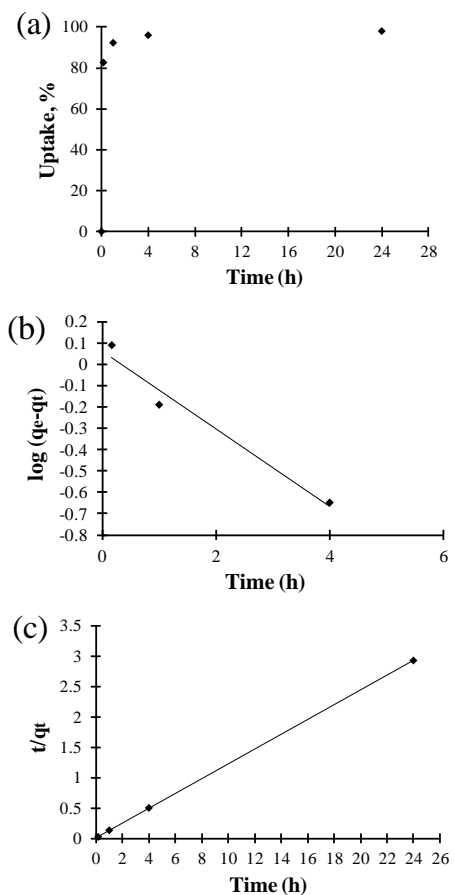


Figure 3. (a) Effect of time on uranium sorption performance of Dowex 2x8 anion exchanger resin, (b) adsorption kinetic curve of pseudo first-order model and (c) linearized experimental data of pseudo second-order kinetics model plot

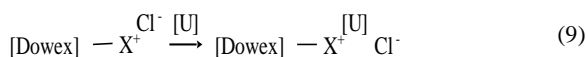
TABLE 2. Kinetic parameters for uranium removal using Dowex 2x8 anion exchanger resin

	k_1 (min ⁻¹)	0.419
Pseudo first-order	q_e (mmol g ⁻¹)	0.861
	R^2	0.9694
	k_2 (g mmol ⁻¹ min ⁻¹)	1.548
Pseudo second-order	q_e (mmol g ⁻¹)	8.203
	R^2	1

The problem that often seems to be challenging is that most of ions existing in the solution have trivial or strong tend to the adsorbent. Hence, the effect of interfering ions on uranium sorption, a challenge which must be met, was evaluated in the presence of sulfate (0-2655 mg l⁻¹), nitrate (0-1000 mg l⁻¹), chloride (0-2527 mg l⁻¹), fluoride (0-4842 mg l⁻¹), and phosphate (0-300 mg l⁻¹) ions.

It can be seen from Figure 4 that, chloride and nitrate ions severely affect uranium uptake using Dowex 2x8 resin rather than the others. The result additionally showed that the presence of high amounts of sulfate, fluoride, and phosphate ions had no interference for uranium sorption in contrary with Cl and NO₃ ions. Having considered the negative impact of Cl and NO₃ ions on the resin performance, it is recommended that before using Dowex 2x8 for uranium sorption, a relevant process for elimination of those ions from the solution should be performed.

3. 5. Anion Exchange Mechanism As a result of inauguration of Dowex 2x8 resin into uranium solution, the following exchange reaction can happen between uranyl ions and the active sites on the surface of the resin (Equation (9)):



The result showed that (based on Figure 4 data) the presence of some anions in the solution can decrease the adsorption efficiency. This effect can be attributed to the formation of various ionic complex between uranium and those anions contents which inhibit the sequestering uranium complexes on the active sites of the resin. As another possible explanation for uranium uptake by Dowex 2x8, we mention the space inhibition of uranium molecule that can play a negative role in sorption process. Under such condition, uranium embedded among the resin's active sites' dentates and they hold uranium complexes in the matrix of resin (Figure 5). Adversely, this phenomenon can attribute to the decrease of uranium removal from solution. As there are other anions such as sulfate, fluoride, chloride, nitrate, and phosphate in the solution, by formation of uranium complexes, they have more opportunity to be adsorbed on resin active sites.

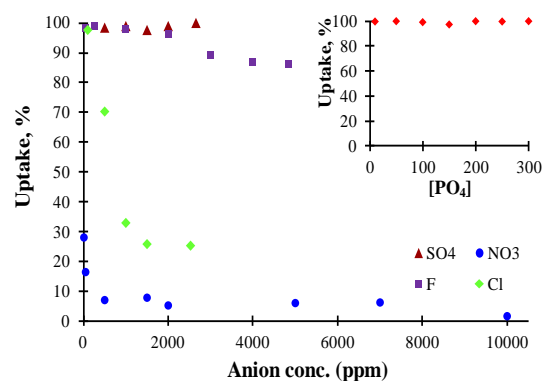


Figure 4 . Effect of presence of various anionic species on uranium (VI) sorption using Dowex 2x8 [sulfate (0-2655 mg l⁻¹), nitrate (0-1000 mg l⁻¹), chloride (0-2527 mg l⁻¹), fluoride (0-4842 mg l⁻¹), phosphate (0-300 mg l⁻¹)]

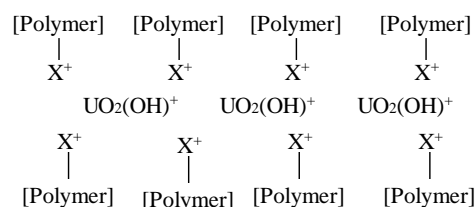


Figure 5. Uranium trapping between resin's active site's dentate

4. CONCLUSION

Briefly, current work investigates affinity of Dowex 2x8 anion exchanger resin towards sorption of uranium ions in the presence of sulfate ions. During the sorption experiments, the operational parameters influencing on U(VI) adsorption onto the resin, including pH, initial concentration, contact time and anion effect were studied. It was disclosed that the sorption strongly occurs at pH 3 and it will remain constant in the afterward pH quantities. To be able to understand the observed phenomena, we employed isotherm and kinetic models to depict the adsorption mathematically. Isotherm study showed that the Freundlich binding model was consistent with experimental data compared to the Langmuir model, indicating the sorption process was influenced by the formation of monolayer coverage of the sorbate (uranium) on the homogenous sorbent surface. Pseudo second-order kinetic model better estimated the result than the pseudo first-order kinetic model. This revealed that the concentration of both sorbate (uranium) and sorbent (Dowex 2x8) are involved in the rate-limiting step of the sorption process. Anions effect including sulfate, nitrate, chloride, fluoride, and phosphate ions on uranium sorption using the Dowex 2x8 resin was investigated. The result showed that the presence of anionic elements such as fluoride, phosphate and sulfate did not affect the

sorption process, while the presence of Cl and NO₃ ions had negative impact on uranium removal. Thus, before using Dowex 2x8 for uranium sorption, a relevant process for elimination of those interfering ions from the solution should be done. The maximum adsorption capacity (q_{max}) of the Dowex 2x8 anion exchanger resin was found to be 303.03 mmol g⁻¹.

5. REFERENCES

1. Taylor, D.M. and Taylor, S.K., "Environmental uranium and human health", *Reviews on Environmental Health*, Vol. 12, No. 3, (1997), 147-158.
2. Shafeeq, A., Muhammad, A., Sarfraz, W., Toqeer, A., Rashid, S. and Rafiq, M., "Mercury removal techniques for industrial waste water", *World Acad. Science Engineering and Technology*, Vol. 6, (2012), 12-26.
3. Zhang, F.-S., Nriagu, J.O. and Itoh, H., "Mercury removal from water using activated carbons derived from organic sewage sludge", *Water Research*, Vol. 39, No. 2, (2005), 389-395.
4. Jeon, C. and Park, K.H., "Adsorption and desorption characteristics of mercury (II) ions using aminated chitosan bead", *Water Research*, Vol. 39, No. 16, (2005), 3938-3944.
5. Kabay, N., Demircioglu, M., Yayli, S., Günay, E., Yüksel, M., Sağlam, M. and Streat, M., "Recovery of uranium from phosphoric acid solutions using chelating ion-exchange resins", *Industrial & Engineering Chemistry Research*, Vol. 37, No. 5, (1998), 1983-1990.
6. Ho, C. and Miller, N., "Adsorption of uranyl species from bicarbonate solution onto hematite particles", *Journal of Colloid and Interface Science*, Vol. 110, No. 1, (1986), 165-171.
7. Kowsari, M.R., Sepehrian, H., Mahani, M. and Fasihi, J., "Cobalt (II) adsorption from aqueous solution using alginate-SBA-15 nanocomposite: Kinetic, isotherm, thermodynamic studies and neural network modeling", *Materials Focus*, Vol. 5, No. 2, (2016), 91-99.
8. Kowsari, M. and Sepehrian, H., "Zinc adsorption properties of alginate-SBA-15 nanocomposite", *International Journal of Engineering Transactions A: Basics*, Vol. 28, No. 10, (2015), 1415-1422.
9. Monteagudo, J.M. and Ortiz, M.J., "Removal of inorganic mercury from mine waste water by ion exchange", *Journal of Chemical Technology and Biotechnology*, Vol. 75, No. 9, (2000), 767-772.
10. Jorgensen, T. and Weatherley, L., "Ammonia removal from wastewater by ion exchange in the presence of organic contaminants", *Water Research*, Vol. 37, No. 8, (2003), 1723-1728.
11. Edebali, S. and Pehlivan, E., "Evaluation of amberlite IRA96 and DOWEX 1×8 ion-exchange resins for the removal of Cr (VI) from aqueous solution", *Chemical Engineering Journal*, Vol. 161, No. 1, (2010), 161-166.
12. Alexandratos, S.D., "Ion-exchange resins: A retrospective from industrial and engineering chemistry research", *Industrial & Engineering Chemistry Research*, Vol. 48, No. 1, (2008), 388-398.
13. Wang, Y.-L., Song, L.-J., Zhu, L., Guo, B.-L., Chen, S.-W. and Wu, W.-S., "Removal of uranium (VI) from aqueous solution using iminodiacetic acid derivative functionalized SBA-15 as adsorbents", *Dalton Transactions*, Vol. 43, No. 9, (2014), 3739-3749.
14. Li, Z., Chen, F., Yuan, L., Liu, Y., Zhao, Y., Chai, Z. and Shi, W., "Uranium (VI) adsorption on graphene oxide nanosheets from aqueous solutions", *Chemical Engineering Journal*, Vol. 210, (2012), 539-546.
15. Makeswari, M. and Santhi, T., "Adsorption of cr (vi) from aqueous solutions by using activated carbons prepared from ricinus communis leaves: Binary and ternary systems", *Arabian Journal of Chemistry*, (2014).
16. Gilani, S.L., Najafpour, G.D., Moghadamnia, A. and Kamaruddin, A.H., "Kinetics and isotherm studies of the immobilized lipase on chitosan support", *International Journal of Engineering, Transactions A: Basics*, Vol. 29, No. 10, (2016), 13.
17. Attia, A.A., Khedr, S.A. and Elkholy, S.A., "Adsorption of chromium ion (vi) by acid activated carbon", *Brazilian Journal of Chemical Engineering*, Vol. 27, No. 1, (2010), 183-193.
18. Ahmadi, S.J., Sadjadi, S. and Hosseinpour, M., "Adsorption behavior of toxic metal ions on nano-structured cuo granules", *Separation Science and Technology*, Vol. 47, No. 7, (2012), 1063-1069.
19. Dada, A., Olalekan, A., Olatunya, A. and Dada, O., "Langmuir, freundlich, temkin and dubinin–radushkevich isotherms studies of equilibrium sorption of zn²⁺ onto phosphoric acid modified rice husk", *Journal of Applied Chemistry*, Vol. 3, No. 1, (2012), 38-45.
20. Vijaya, Y., Popuri, S.R., Boddu, V.M. and Krishnaiah, A., "Modified chitosan and calcium alginate biopolymer sorbents for removal of nickel (II) through adsorption", *Carbohydrate polymers*, Vol. 72, No. 2, (2008), 261-271.
21. Rao, M.M., Rao, G.C., Seshiah, K., Choudary, N. and Wang, M., "Activated carbon from ceiba pentandra hulls, an agricultural waste, as an adsorbent in the removal of lead and zinc from aqueous solutions", *Waste Management*, Vol. 28, No. 5, (2008), 849-858.

Uranium Removal from Aqueous Solution Using Ion-exchange Resin DOWEX[®] 2x8 in the Presence of Sulfate Anions

M. R. Kosari^a, H. Sepehrian^b, R. Salamat Ahangari^c

^a Department of Energy Engineering, Sharif University of Technology, Azadi Avenue, Tehran, Iran.

^b Nuclear Science and Technology Research Institute, End of North Karegar Avenue, Tehran, Iran

^c Department of Chemistry, Faculty of Basic Science, Azarbaijan Shahid Madani University, Tabriz, Iran

P A P E R I N F O

چکیده

Paper history:

Received 01 June 2016

Received in revised form 04 October 2016

Accepted 11 November 2016

Keywords:

Dowex 2x8

Anion Exchanger Resin

Anion Effect

Adsorption

Uranium

Heavy Metals.

مطالعه ی حاضر بر روی جذب یون های اورانیوم (VI) با استفاده از یک رزین تبادلگر آنیونی، Dowex[®]2x8، در حضور یونهای سولفات انجام شده است. ارزیابی پارامترهای عملیاتی شامل زمان تماس، pH، غلظت اولیه یون های اورانیوم و حضور گونه های آنیونی مختلف (شامل فسفات، سولفات، کلراید، فلوئور و نیترات) بر روی رفتار جذبی Dowex 2x8 در محلول ارزیابی شد. تحلیل ایزوترم با استفاده از مدل های همبستگی لانگمویر و فروندلیچ به منظور مشخص کردن رابطه ی بین یون های اورانیوم جذب شده بر روی رزین و غلظت یون های اورانیوم باقی مانده در فاز آبی انجام گردید. برای آشکار کردن مرحله ی محدود کننده ی نرخ جذب، مدل سازی سینتیکی با استفاده از معادلات شبه درجه اول و شبه درجه دوم نیز انجام گردید. بیشترین میزان ظرفیت جذب برای رزین تبادلگر آنیونی Dowex 2x8 در فرآیند حذف اورانیوم برابر با $3.3/0.3 \text{ mmol.g}^{-1}$ بود. همچنین مشخص گردید که حضور گونه های آنیونی متفاوت در محلول می تواند جذب یون های اورانیوم توسط رزین تبادلگر آنیونی Dowex 2x8 را تحت تاثیر قرار دهد.

doi: 10.5829/idosi.ije.2016.29.12c.06